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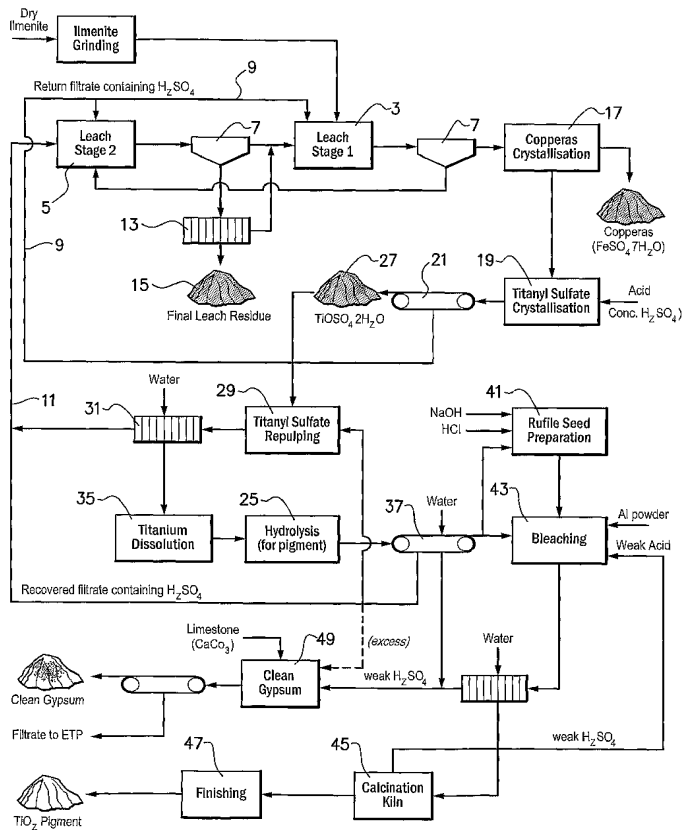
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[Continued on next page]

**(54) Title:** OPERATING TITANIUM PRECIPITATION PROCESS



**(57) Abstract:** A sulfate process for producing titania from a titaniferous material as disclosed. The process is characterized by a particular step of precipitating titanyl sulfate.



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OPERATING TITANIUM PRECIPITATION PROCESS

The present invention relates to a process for producing titania from a titaniferous material.

The term "titaniferous" material is understood herein to mean any titanium-containing material, including by way of example ores, ore concentrates, and titaniferous slags.

The present invention relates particularly to the sulfate process for producing titania from titaniferous material.

International application PCT/AU2004/001421 in the name of the applicant describes an invention of a sulfate process made by the applicant. The disclosure in the International application is incorporated herein by cross-reference.

In general terms, the present invention provides a sulfate process for producing titania from a titaniferous material (such as ilmenite) of the type which includes the steps of:

(a) leaching the solid titaniferous material with a leach solution containing sulfuric acid and forming a process solution that includes an acidic solution of titanyl sulfate ( $\text{TiOSO}_4$ ) and iron sulfate ( $\text{FeSO}_4$ );

(b) separating the process solution and a residual solid phase from the leach step (a);

(c) precipitating titanyl sulfate from the process solution from step (b);

(d) separating the precipitated titanyl sulfate

- 2 -

from the process solution;

(e) treating the precipitated titanyl sulfate and producing a solution containing titanyl sulfate;

5

(f) hydrolysing the titanyl sulfate in the solution and forming a solid phase containing hydrated titanium oxides and a liquid phase;

10

(g) separating the solid phase containing hydrated titanium oxides and the liquid phase;

(h) calcining the solid phase from step (g) and forming titania; and

15

(i) removing iron sulfate from the process solution from step (b) and/or the depleted process solution from step (d).

20

The term "hydrated titanium oxides" is understood herein to include, by way of example, compounds that have the formula  $\text{TiO}_2 \cdot 2\text{H}_2\text{O}$  and  $\text{TiO}_2 \cdot \text{H}_2\text{O}$ .

25

In addition, the term "hydrated titanium oxides" is understood herein to include compounds that are described in technical literature as titanium hydroxide ( $\text{Ti}(\text{OH})_4$ ).

30

It is also noted at this point that acid concentrations mentioned hereinafter are understood herein as being determined by titration of an oxalate buffered solution sample with sodium hydroxide solution to an end point of pH 7.

35

It is also noted at this point that concentrations of metals mentioned hereinafter are understood herein as being determined by ICP (all metals)

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or by titration (in the cases of Ti and Fe - ferrous and ferric).

As is indicated in the above-mentioned  
5 International application, US patent 3,760,058 in the name Langmesser et al (assigned to Farbenfabriken Bayer AK) discloses a part of the above-described process.

The reference herein to the Bayer US patent is  
10 not to be taken as an indication that the disclosure in the patent is part of the common general knowledge of persons skilled in the field of the invention.

Preferably the process further includes supplying  
15 the separated process solution from step (d) and/or the separated liquid phase from step (g) to leach step (a).

The applicant has carried out further research  
work on the process since the priority date of 17 October  
20 2003 of the International application and has identified a number of features that are not disclosed in the above-mentioned International application that are important, separately and in combination, in order to operate the process effectively and that form the basis of the present  
25 invention.

The present invention is based on features of the  
step (c) of precipitating titanyl sulfate that are  
described hereinafter that were identified in the further  
30 research work

Other features of the above-described process  
that were identified in the further research work are  
described in Australian provisional application 2005901749  
35 in the name of the applicant and the disclosure in this specification is incorporated herein by cross-reference.

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Titanyl sulfate precipitation step (c) may be carried out on a continuous basis in a single reactor or in a plurality of reactors in series and/or in parallel.

5           Step (c) may also be carried out on a batch basis in a single reactor or in a plurality of reactors.

10           In the above described research work the applicant has found that it is preferred that the reactor or reactors include a vertically extending open-ended draft tube that divides the or each reactor into an inner chamber (defined by the draft tube) and an outer chamber and a stirring means in the tank.

15           Preferably step (c) includes supplying process solution containing dissolved titanyl sulfate from steps (a) and/or (i) to the draft tube reactor or reactors and circulating the process solution successively through the draft tube and the outer chamber of the reactor or  
20           reactors for a sufficient period of time to allow precipitation of titanyl sulfate from solution in the process solution.

25           Step (c) may be carried out on a continuous basis. In this embodiment step (c) may include supplying process solution to the reactor or reactors on a continuous basis and discharging process solution containing suspended precipitated titanyl sulfate from the reactor or reactors on a continuous basis.

30           Step (c) may also be carried out on a batch basis. In this embodiment step (c) may include filling the reactor or reactors with process solution, circulating process solution within the reactor or reactors, partially  
35           or completely discharging process solution containing suspended precipitated titanyl sulfate from the reactor or reactors, and repeating the above steps.

- 5 -

In addition, the applicant has also found that in one embodiment it is preferable that there be a relatively low concentration of titanium in the process solution during step (c).

Preferably the titanium concentration is less than 25g/L. The low concentration of titanium is important to keep the resulting slurry "fluid" during titanyl precipitation in step (c) and in subsequent downstream processing of the slurry.

Furthermore, the applicant has also found that in one embodiment it is preferable to control the process so that there is less than 10 g/L difference between concentration of titanium in the process solution supplied to step (c) and process solution containing suspended precipitated titanyl sulfate discharged from step (c).

The applicant has also found that titanyl sulfate precipitation is essentially self-seeded by virtue of precipitated titanyl sulfate circulating within the reactor or reactors.

Seed may be added to the reactor or reactors as fresh seed and/or as recycled seed.

Furthermore, the applicant has also found that in one embodiment it is preferable that the solids loading in process solution containing suspended precipitated titanyl sulfate discharged from step (c) be less than 10% solids by weight, more preferably 3-4% solids by weight.

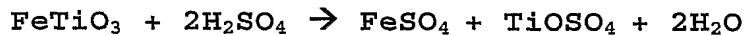
The precipitated titanyl sulfate is typically in the form of 100 x 1 micron needles that are difficult to handle and do not settle readily from the leach liquor. The applicant has found that there are significant

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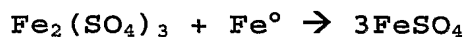
materials handling issues with solids loadings greater than 10% solids by weight.

5 The process of the present invention includes the following typical reactions.

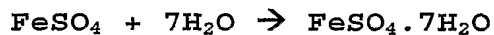
Leaching:



10 Ferric reduction:

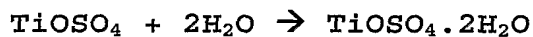


Ferrous sulfate crystallisation:

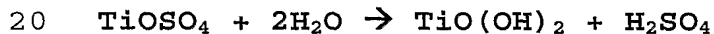


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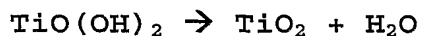
Titanyl sulfate precipitation:



Hydrolysis:



Calcination:



25 The applicant has carried out experimental work on a laboratory scale and a pilot plant scale in relation to the above-described process.

30 The improved sulfate process of the present invention is now described further by way of example only with reference to the accompanying flow sheet.

The flowsheet includes the following main steps:

35 (a) leach;

(b) ferrous sulfate crystallisation;



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(c) titanyl sulfate crystallisation;

(d) titanium dissolution;

5

(e) hydrolysis for pigment;

(f) rutile seed preparation;

10

(g) bleaching,

(h) calcination; and

(i) finishing.

15

Each of the above steps is described hereinafter in turn.

(a) Leach Step

20

The leach step includes two leach stages 1 and 2 carried out in separate tanks 3,5.

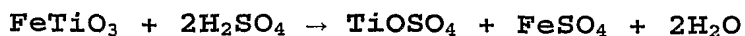
25

Each leach stage is carried out in a single tank 3,5 as indicated in the flowsheet or in multiple tanks (not shown) arranged in series.

30

The leach stages 1 and 2 may be a fully counter-current or may be co-current with fresh return filtrate and/or wash filtrates being added to both leach stages.

The chemistry of the leach step is:



35

Leaching is carried out at a controlled acidity of 450 g/L ( $\pm 25$ g/L)  $\text{H}_2\text{SO}_4$  in each stage. Under these

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conditions about 80% leaching takes place in two leach stages, each of about 12 hours residence time.

5 The leaching temperature is typically 110°C in each stage, which is less than the solution boiling point. The temperature is not controlled, but sufficient heat is generated during leaching to keep the slurry at about 110°C. Some top-up steam may be required for start up.

10 One option is to use scrap iron addition into the leach tanks 3, 5. This has been found to increase leach kinetics significantly. Some reductant is required to convert ferric sulfate to ferrous sulfate to allow all iron to exit in the form of  $\text{FeSO}_4$  crystals.

15 The leach tanks 3, 5 are simple stirred tanks, each of which operates with an overflow to a thickener 7. Fibre-reinforced plastic (FRP) is suitable for wetted parts. Other suitable materials are acid bricks and  
20 tiles.

The leach tanks 3, 5 are operated with gentle stirring so that the residence time of solids in the tanks is longer than the residence time of liquor in the tanks.

25 The leach slurries discharged from the tanks 3, 5 are thickened in conventional thickeners 7. The settling rate is high for partly reacted ilmenite. Flocculation is possible. Underflow densities exceeding 60% are feasible,  
30 but lower solids loadings may be required to ensure pumpability.

The solids loading in the leach step is controlled to give a process solution of about 40 g/L Ti,  
35 90-100 g/L Fe and 400-450 g/L acid that leaves the leach step as overflow from the downstream thickener 7. These are the preferred concentrations of Fe and Ti without

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having ferrous sulfate or titanyl sulfate crystallise out prematurely.

Ilmenite is added dry to the first leach tank 3.

5

To control the acidity to 450 g/L ( $\pm 25$  g/L) return filtrate from the titanyl sulfate crystallisation step 19 discussed hereinafter is supplied via line 9 to the tanks 3, 5 and/or additional sulfuric acid is metered into the tanks 3, 5. In situations where there are multiple tanks 3, 5 in each stage, most of the acid is added to the first two tanks 3, 5 in each stage. In practice, the acidity in later tanks may be uncontrolled.

10

15

Thickener underflow from the thickener 7 of the first leach stage is pumped to the leach tank 5 of the second leach stage.

20

Some recycled acid at about 350 g/L ( $\pm 25$  g/L)  $H_2SO_4$ , which is a filtrate from a filtration step 37 downstream of a hydrolysis step 25 described hereinafter, is also pumped via line 11 to the leach tank 5.

25

Titanyl sulfate crystallisation filtrate produced in a filtration step 31 described hereinafter is also added via line 11 to the second tank 5 to maintain the acidity at 450 g/L ( $\pm 25$  g/L).

30

Leaching is about 50-60% in the first stage rising to about 80% overall by the end of the second stage. Higher extractions are feasible with further leaching.

35

The second stage leach slurry that is discharged from the leach tank 5 is thickened in the thickener 7.

In a full counter-current operation the second

- 10 -

stage overflow from the thickener 7 is pumped to the first stage leach tank 3. In a co-current circuit the solids loading is higher in both stages so that the target of 40 g/L Ti is reached in the final process solution.

5

Second stage leach residue is filtered via filter 13 and the resilient filter cake is suspended in recycled water. Limestone and lime are added to raise the pH to 7-8, and the slurry is pumped to tailings 15.

10

The process solution contained in the (unwashed) filter cake that is supplied to tailings 15 represents the major outlet for a number of minor elements, such as Cr and Zn.

15

Low acidity in the leach stages can cause the premature hydrolysis and precipitation of  $\text{TiO}(\text{OH})_2$ . Typically this becomes significant below about 425 g/L  $\text{H}_2\text{SO}_4$ . Above 450 g/L  $\text{H}_2\text{SO}_4$  it is likewise possible to prematurely crystallise out titanyl sulfate dihydrate  $\text{TiOSO}_4 \cdot 2\text{H}_2\text{O}$ .

20

#### (b) Ferrous Sulfate Crystallisation Step

25

Almost all iron in solution eventually leaves the circuit as green crystals of ferrous sulfate,  $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$ , in a ferrous sulfate crystallization step 17.

30

Significant water is also rejected from the process in the ferrous sulfate, also known as "copperas". This allows recovery and recycling of medium strength acid from the hydrolysis step, leading to a much lower overall acid consumption per tonne of  $\text{TiO}_2$  product.

35

In the ferrous sulfate crystallization step 17, hot process solution discharged as the overflow from the downstream thickener 7 of the leach step is firstly cooled

- 11 -

to about 60°C in a heat exchanger (not shown) by heat exchange with process solution that has been discharged from a downstream crystallization tank (not shown).

5           The cooled pregnant process solution is then evaporatively cooled to about 20°C. This causes ferrous sulfate to crystallise out in the tank. The cooled process solution at this stage contains about 40 g/L Fe and 55 g/L Ti. The Ti concentration rises due to the  
10 lower volume of the cooled process solution.

Removal of water by evaporation may be included to give a further water credit, allowing recovery of more weak acid.

15           The ferrous sulfate crystals may be separated from the process solution by a conventional centrifuge (not shown) or by a belt filter (not shown).

20           Some washing may be possible, but the high solubility of the crystals means that washing should be minimised if possible.

25           The ferrous sulfate may be sold directly or converted to another saleable product.

30           Although 40 g/L Fe remains in solution, the iron is recirculated and eventually returns to the ferrous sulfate crystallization step 17. The ferrous sulfate crystals therefore are essentially the only point of exit for iron from the circuit.

Mn, Al and Mg are minor elements which exit the circuit primarily with the ferrous sulfate crystals.

35           Finally, the cold process solution that is discharged from the ferrous sulfate crystallization step

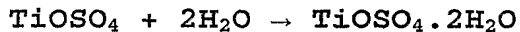
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17 is partially reheated by cross flow heat exchanging against incoming hot process solution supplied to the step 17.

5 (c) Titanyl Sulfate Precipitation Step

Fresh 98% sulfuric acid that is required for leaching ilmenite is not added in the leach stages of the leach step. Instead, the acid is added in a titanyl sulfate precipitation step, generally identified by the numeral 19.

The acid causes titanium to precipitate out of the process solution as titanyl sulfate dihydrate,  $\text{TiOSO}_4 \cdot 2\text{H}_2\text{O}$  and form a slurry in accordance with the following reaction:



The actual mechanism of precipitation is not clear.

The preferred operating temperature in the titanyl sulfate precipitation step is 110°C. Precipitation is very slow at less than 90°C.

Precipitation is self seeding - the kinetics of crystallisation are accelerated by the presence of the product crystals.

30

The solids have a long needle-like shape (typically 1µm width by 100µm long). The needle-like morphology causes significant rheology problems in the titanyl sulfate precipitation step. Quite low solids loadings can result in thick porridge-like slurries which resist pumping and agitation.

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In one particular embodiment the precipitation tank (or one or more than one of the precipitation tanks in situations where there are multiple tanks) has an upstanding draft tube that has an upper inlet and a lower outlet and the draft tube is located to divide the tank into an outer chamber and a central cylindrical chamber. The assembly also includes an impeller to help circulation of the slurry. The slurry flows through the draft tube and the outer chamber in the tank.

To keep the slurry in a fluid state a recycle of filtrate may be used.

The solids in the slurry that is discharged from the precipitation tank or tanks are separated from the slurry by filtration. Filtration may be by a belt filter 21 shown in the flowsheet. However, maintaining the temperature of the filtrate probably requires pressure filtration.

Some washing of the solids in the filter cake on the filter 21 by recycled acid from the hydrolysis step described hereinafter may be carried out as this improves purity of the high strength Ti solution going to hydrolysis.

The acid washed  $\text{TiOSO}_4 \cdot 2\text{H}_2\text{O}$  filter cake is a stable solid product and offers a convenient breakpoint in the flowsheet. The filter cake may be stock-piled as indicated by the numeral 27. Temporary storage of the acid washed crystals offers useful buffer capacity, and makes the process more robust.

The filtrate contains about 700 g/L  $\text{H}_2\text{SO}_4$  (roughly 50% w/v) plus 10 g/L Ti and 40 g/L Fe. Some is recycled to the titanyl sulfate precipitation stage tank 19. The rest is sent to the leach stages via line 9,

- 14 -

where it is used to control the acidity to 450 g/L  $\text{H}_2\text{SO}_4$  in the leach slurry.

Thickening before filtration is not used due to  
5 the needle-like solids, which do not compact readily under gravity.

(d) Titanium Dissolution

10 The acid washed filter cake from the stockpile 27 is re-pulped in a 30%  $\text{H}_2\text{SO}_4$  solution in a re-pulping step 29 and is then is pumped to a filter 31. The resultant slurry has an acid concentration of the order of 400 g/L.

15 The filter cake on the filter 31 may be washed with hydrolysis filtrate to remove remaining entrained leach liquor.

Finally, a carefully controlled water wash is  
20 used to displace all the remaining acid in the filter cake on the filter 31. Reducing the acid concentration to below 200 g/L destabilises the solids, leading to ultimate dissolution of the solids. Cake squeezing and/or air blowing is then used to control the moisture content of  
25 the cake. About 5 g/L Ti reports to the wash filtrate, which is recycled via line 11 to the leach stages.

As described above, these washing steps may be applied to the initial filtration step to eliminate the  
30 need to re-pulp and re-filter the solids. However, in doing so the ability to store an intermediate filter cake is lost and the process is less robust.

The water washed filter cake discharged from the  
35 filter 31 is added to a stirred tank 35. Over a period of about 2 hours at 60°C the cake dissolves into a high strength Ti solution. Lower temperatures can also be



- 15 -

used, although the dissolution time may be longer than 2 hours.

5       The target concentration is 150 g/L Ti (= 250 g/L "TiO<sub>2</sub>"). Concentrations exceeding 200 g/L Ti have been produced in laboratory and pilot plant work. However, 150 g/L or above is suitable for conventional pigment hydrolysis.

10       The dissolution process preferably requires less than 100 g/L acid in the solution contained within the filter cake to ensure that the process goes to completion. If most or all acid is washed out the free acid content of the high strength solution is quite low. In pigment  
15 industry terms the acid to titania (A/T) ratio is usually about 1.3 (the theoretical minimum is 1.225 at zero acidity).

20       The product high strength solution produced in the stirred tank 35 is filtered through a filter cartridge (not shown) to remove siliceous and other fine particulate matter.

25       Unlike normal metal sulfates, the TiOSO<sub>4</sub>.2H<sub>2</sub>O in the filter cake does not immediately dissolve in water. Also its solubility in >20% H<sub>2</sub>SO<sub>4</sub> is quite low. This suggests the dissolution process is not strictly dissolution. The remarkable solubility of Ti at low acidity (>200 g/L Ti) compared to 20% H<sub>2</sub>SO<sub>4</sub> (~5 g/L Ti)  
30 favours this view.

#### (e) Hydrolysis Step

35       The filtered high strength Ti process solution is suitable for all conventional pigment hydrolysis processes.

- 16 -

It also may be used for continuous or batch precipitation of coarse high purity  $\text{TiO}(\text{OH})_2$ .

5 The pigment hydrolysis processes are typically batch processes due to critical need to control particle size.

10 Feed solution to the pigment hydrolysis step is pretreated to generate about 2 g/L of  $\text{Ti}^{3+}$  in the solution by conventional means. The  $\text{Ti}^{3+}$  protects against oxidation of any iron to  $\text{Fe}^{3+}$ , which coprecipitates with the Ti and imparts undesirable colour to the pigment.

15 The process solution is then adjusted with acid to an A/T ratio suitable for pigmentary hydrolysis, using either concentrated  $\text{H}_2\text{SO}_4$  or preferably the hydrolysis filtrate. The A/T ratio is a key process parameter. A/T ratio is:

20 
$$[\text{Free acid} + \text{bound acid in TiSO}_4] \div [\text{TiO}_2]$$

All parameters are expressed in g/L.

25 In practice the [Free acid + bound acid in  $\text{TiOSO}_4$ ] concentration is measured by a simple titration to pH 7 with sodium hydroxide solution, and the  $[\text{TiO}_2]$  g/L is  $\text{Ti g/L} \div 0.6$ .

30 In one example of commercial practice, the hydrolysis is carried out by preheating a heel of water, typically 10-20% of the volume of feed solution, to about  $96^\circ\text{C}$ .

35 The process solution is also preheated to about  $96^\circ\text{C}$ , then is pumped across to the batch hydrolysis tank over a fixed time period.

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The hydrolysis tank 25 is equipped with steam heating and a gate type rake stirrer, which operates at low rpm. Preferably the steam heating is indirect so that the filtrate is not diluted by condensate.

5

The initial few seconds of pumping cause the precipitation of very fine  $\text{TiO}(\text{OH})_2$  particles, which cause a milky aspect for about 30 seconds, then appear to redissolve. In practice the fine particles are colloidal nuclei which control the size of both the resulting precipitate and the crystal size in the calciner discharge. Control of this step is therefore key to preparing good pigment.

10

15

After all process solution is pumped across or dropped in from a header tank, the slurry temperature is carefully heated to the boiling point (typically at  $1^\circ\text{C}/\text{minute}$ ).

20

The slurry is then boiled for about 5 hours, by which time the Ti remaining in solution has been lowered to about 5 g/L.

25

The slurry discharged from the hydrolysis tank 25 is filtered and washed with water on a belt filter 37 and produces a  $\text{TiO}(\text{OH})_2$  filter cake and a filtrate.

30

There are no special requirements for filtration as the particle size has already been established. A range of filters are used across the industry. The particles naturally floc together and the filtration rate is fast enough that vacuum filtration may be used. The filter cake contains about 55% w/w of water.

35

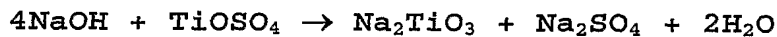
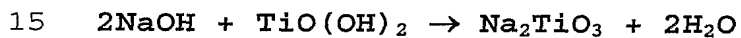
The filtrate from the filter 37 contains 350-450 g/L  $\text{H}_2\text{SO}_4$ . This is returned via line 11 to the leach step for slurrying ilmenite and/or first stage thickener

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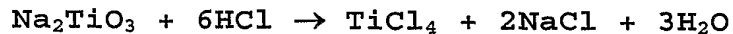
underflow. The acid units thereby are used to leach ilmenite. Recycling this acid is limited by the overall circuit water balance, and is favoured by higher acidity (ie. a lower volume equates to the higher acidity). Any excess is sent to a clean gypsum plant 49.

(f) Rutile Seed Preparation Step

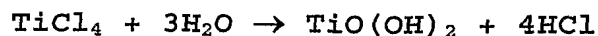
In one example of commercial practice, rutile seed is made in a rutile seed preparation step 41 by reacting some  $\text{TiO}(\text{OH})_2$  filter cake discharged from the belt filter 37 with commercial 50% NaOH solution, for several hours at the boiling point (about  $117^\circ\text{C}$ ):



The  $\text{TiO}(\text{OH})_2$  filter cake contains about 4% S in the form of absorbed basic titanium sulfates. The resulting sodium titanate is filtered and washed well to completely remove sulfate. The washed cake is then mixed with a carefully controlled amount of commercial 35% HCl to produce a solution of  $\text{TiCl}_4$ ;



The solution is then boiled to generate ultrafine  $\text{TiO}(\text{OH})_2$  particles:



The resulting slurry contains about 100 g/L  $\text{TiO}_2$  in the rutile form. It may be used directly if the downstream flowsheet can tolerate Cl or it can be decantation washed to remove the NaCl.

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## (g) Bleaching Step

The  $\text{Ti}(\text{OH})_2$  filter cake that is discharged from the belt filter 37 and is not used to make rutile seed is re-pulped with clean  $\text{H}_2\text{SO}_4$  solution in a bleaching step 43. Al or Zn dust is added to reductively leach out chromophores such as Fe, Cr, Mn and V, which otherwise would reduce the whiteness of the final pigment.

The bleach step typically takes place at  $80^\circ\text{C}$ . The rutile seed slurry is added at this point in a carefully controlled amount (e.g.  $4.0 \pm 0.1$  % w/w). The bleached slurry is filtered and washed.

The  $\text{TiO}(\text{OH})_2$  filter cake, which has a sulfur content of about 2%, is mixed with a number of additives. These may be added as water solutions, or solids. The additives may include 0.2%  $\text{K}_2\text{O}$  as  $\text{K}_2\text{SO}_4$ , 0.6% ZnO as  $\text{ZnSO}_4$  and 0.3%  $\text{P}_2\text{O}_5$  as  $\text{H}_3\text{PO}_4$ .

The additives control development of the rutile crystals during calcination, such that the crystal size is  $0.27 \pm 0.03\mu\text{m}$ , rutilisation is  $98.5 \pm 0.5\%$ , the crystals have a lenticular shape and are not sintered together.

In addition to the above-described steps, the process flowsheet also includes the steps of: calcination 45, finishing 47, and, if required, clean gypsum production 49. These steps are conventional steps.

Many modifications may be made to the process flowsheet described above without departing from the spirit and scope of the present invention.

By way of example, as an alternative to pigment production, the process is able to produce coarse high purity titania that can be used, for example, as a

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feedstock for electrochemical reduction to produce titanium metal and alloys. Hydrolysis may be carried out continuously in this option. Several simple stirred tanks may be used in a cascade arrangement. Hydrolysis may be carried out at boiling point using steam heating, preferably indirect. Seeding is carried out by recycling thickener underflow to the first tank. This allows the slurry residence time to be 8-12 hours and generates a particle size  $d_{50}$  of about 20 microns. Thickening gives a dense slurry of about 30% solids by weight, which may be vacuum filtered and washed. Bleaching may be carried out per the pigment process, if required. No rutile or chemical seeds are used. Calcination only requires a temperature of the order of 900°C for about 1 hour.

The present invention is described further with reference to the following examples.

Within these examples where 'free  $H_2SO_4$ ' has been referred to, this has been determined by titration of an oxalate buffered solution sample with sodium hydroxide solution to an end point of pH 7.

#### Example 1

This example describes a first stage of batch leaching.

A solution (300L) containing 3.0 g/L Ti, 11.2 g/L  $Fe^{2+}$ , 3.0 g/L  $Fe^{3+}$ , and 716 g/L free  $H_2SO_4$  was heated in a stirred, baffled vessel. Once the liquor had reached 110°C, 79.6 kg of ilmenite concentrate containing 25.9% FeO, 19.3%  $Fe_2O_3$  and 50.4%  $TiO_2$ , which had previously been ground in a ball mill to 80% less than 38  $\mu m$ , was introduced into the reaction vessel. Six 10 mm diameter mild steel rods were suspended in the reactor such that about 200 mm of the rods extended below the solution

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level. The mixture was allowed to react at 110°C for 3 hours, after which the temperature was allowed to fall steadily to 80°C over the next 3 hours. The resulting slurry was filtered through a recessed plate filter and the cake was washed with fresh water. The filtrate contained 47 g/L Ti, 55 g/L  $\text{Fe}^{2+}$ , 17 g/L  $\text{Fe}^{3+}$ , 618 g/L free  $\text{H}_2\text{SO}_4$ , and had a specific gravity of 1.637 g/cm<sup>3</sup>. The weight of the washed filter cake was 39kg with a moisture content of 16.9%. The washed filter cake was assayed on a dry weight basis and was found to contain 15.3% FeO, 24.4%  $\text{Fe}_2\text{O}_3$  and 48.7%  $\text{TiO}_2$ .

Based on the weights and compositions of the ilmenites and cake, 60.6% of the  $\text{TiO}_2$  in the ilmenite dissolved during the leach process.

#### Example 2

This example describes a second stage of leaching using the first stage leach residue.

A solution (273 L) containing 3.6 g/L Ti, 6.1 g/L  $\text{Fe}^{2+}$ , 2.4 g/L  $\text{Fe}^{3+}$ , and 711 g/L free  $\text{H}_2\text{SO}_4$  was heated in a stirred, baffled vessel. Once the liquor had reached 110°C, 130 kg of wet cake prepared as described in Example 1, having a moisture content of 18.6% and containing 17.0% FeO, 22.7%  $\text{Fe}_2\text{O}_3$  and 49.4%  $\text{TiO}_2$ , was introduced into the reaction vessel. Six 10 mm diameter mild steel rods were suspended in the reactor such that about 200 mm of the rods extended below the solution level. The mixture was allowed to react at 110°C for 3 hours, after which the temperature was allowed to fall steadily to 80°C over the next 3 hours. The resulting slurry was filtered through a recessed plate filter and the cake was washed with fresh water. The filtrate contained 46 g/L Ti, 38 g/L  $\text{Fe}^{2+}$ , 20 g/L  $\text{Fe}^{3+}$ , 513 g/L free  $\text{H}_2\text{SO}_4$ , and had a specific gravity of 1.553 g/cm<sup>3</sup>. The weight of the washed filter cake was 86

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kg with a moisture content of 26.2%. The washed filter cake was assayed on a dry weight basis and was found to contain 13.3% FeO, 22.7% Fe<sub>2</sub>O<sub>3</sub> and 49.7% TiO<sub>2</sub>.

5           Based on the weights and compositions of the feed and product and cakes, 39.7% of the TiO<sub>2</sub> in the feed cake dissolved during the leach process.

### Example 3

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This example describes the reduction and removal of Fe<sup>3+</sup> from the solution produced as described in Examples 1-2.

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A 5 L baffled glass reactor fitted with an 80 mm Rushton 6 turbine agitator was filled with 4 L of a solution containing 13.2 g/L Fe<sup>3+</sup>, 38.5 g/L Fe<sup>2+</sup>, 505 g/L free H<sub>2</sub>SO<sub>4</sub> and 40 g/L Ti. The agitation rate was set at 500 rpm. The reactor was temperature controlled to 50°C. On reaching this temperature a pump was used to recirculate the solution at 100 mL/min from the glass vessel, and through a 4 L fibre reinforced plastic (FRP) vessel containing a single 150 mm x 150 mm x 150 mm compressed bale of commercial detinned scrap steel. The solution was introduced to the bottom of the FRP vessel and flowed up through the scrap and overflowed via gravity back into the glass reactor. The bale of scrap was height adjusted to be fully submerged below the level of the solution in the FRP vessel. After recirculating the solution for 45 min it was found that all Fe<sup>3+</sup> had been consumed. After 60 minutes the pump was turned off and the bale of scrap removed, whereupon it was found the solution contained 0 g/L Fe<sup>3+</sup>, 93 g/L Fe<sup>2+</sup> and 8.5 g/L Ti<sup>3+</sup>.

35   Example 4

This example shows that ferrous sulfate may be



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batch precipitated from an ilmenite leach solution.

An ilmenite leach solution containing 0.1 g/L  $\text{Fe}^{3+}$ , 98.2 g/L  $\text{Fe}^{2+}$ , 48 g/L Ti and 399 g/L free  $\text{H}_2\text{SO}_4$ , prepared in the manner described in Example 3, was placed in a beaker and cooled overnight. Green ferrous sulfate heptahydrate crystals with composition 18.5% Fe, 10.5% S, 0.23% Ti and 0.15% Mn were then recovered from the resulting slurry. The filtrate was assayed and found to contain <1 g/L  $\text{Fe}^{3+}$ , 30.2 g/L  $\text{Fe}^{2+}$  and 539 g/L free  $\text{H}_2\text{SO}_4$ .

#### Example 5

This example shows that titanyl sulfate dihydrate,  $\text{TiOSO}_4 \cdot 2\text{H}_2\text{O}$ , crystals may be batch precipitated from an ilmenite leach solution prepared in the manner of Examples 1-2 by the addition of sulfuric acid, and that a high strength solution suitable for pigment manufacturing may be generated by dissolution of the crystals.

Sulfuric acid (98%, 450 g) was mixed with an ilmenite leach solution (1500 mL) containing 440 g/L free  $\text{H}_2\text{SO}_4$ , 35.4 g/L  $\text{Fe}^{2+}$ , 7.4 g/L  $\text{Fe}^{3+}$  and 29 g/L Ti in a glass reactor equipped with baffles and a Teflon agitator. The resulting solution was heated to 110° C and titanyl sulfate crystals (4 g) were added as seed material. The mixture was stirred at this temperature for a total of 6 hours, during which a thick precipitate formed. The slurry was filtered and the cake was washed with water to give a wet filter cake (238 g). The filtrate contained 16 g/L Ti, 638 g/L  $\text{H}_2\text{SO}_4$  and 48 g/L Fe, of which 6.6 g/L was as  $\text{Fe}^{3+}$ . The filter cake dissolved after 3 hours to produce a titanyl sulfate solution containing 160 g/L Ti and 8.3 g/L Fe.

#### Example 6

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This example describes the continuous precipitation of titanyl sulfate dihydrate,  $\text{TiOSO}_4 \cdot 2\text{H}_2\text{O}$ , crystals using a reactor fitted with a draft tube, followed by vacuum filtration.

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Ilmenite leach solution (603.6 L) prepared as described in Examples 1-2, containing 524.7 g/L free  $\text{H}_2\text{SO}_4$ , 14.5 g/L  $\text{Fe}^{2+}$ , 4.3 g/L  $\text{Fe}^{3+}$  and 41.2 g/L Ti was mixed in an agitated fibreglass reactor with titanyl sulfate filtrate (1043.2 L) containing 637.5 g/L free  $\text{H}_2\text{SO}_4$ , 44.7 g/L  $\text{Fe}^{2+}$ , 12.8 g/L  $\text{Fe}^{3+}$  and 6.1 g/L Ti. Sulfuric acid (98%, 88.3 L) was then added along with titanyl sulfate filter cake (10 kg, 14% w/w solids) and the temperature was raised to 110°C. The reactor was 1.35 m diameter, with 1.3 m solution depth and contained a draft tube to improve mixing and the uniformity of mixing inside the reactor with minimal power input. The draft tube was 0.9 m internal diameter, 0.87 m high and raised 0.25 m from the bottom of the reactor. The reactor was fitted with an axial turbine with diameter of 0.6 m and raised 0.5 m from the floor of the reactor. The turbine operated at 250 rpm. The reactor was allowed to stir at temperature for 12 hours and a sample was taken and filtered. The titanium concentration in the liquor had dropped from an initial combined level of 17.3 g/L to 9.0 g/L. The feed and product pumps were started and set to flowrates of 4.6 L/min to allow for a 4.9 hour residence time with a constant combined feed solution containing 17.5 g/L Ti and 660 g/L  $\text{H}_2\text{SO}_4$ . The precipitator was run continuously this way for 10 hours producing 2742 L of titanyl sulfate slurry. Regular samples were taken from the reactor and filtered and analysed. These filtrate samples gave average concentrations of 7.5 g/L Ti and 611.8 g/L  $\text{H}_2\text{SO}_4$ . The precipitated titanyl sulfate dihydrate ( $\text{TiOSO}_4 \cdot 2\text{H}_2\text{O}$ ) was separated from the slurry using a belt filter, giving approximately 780 kg of filter cake with solids loading 14% w/w.

## Example 7

5 This example demonstrates that titanyl sulfate dihydrate,  $\text{TiOSO}_4 \cdot 2\text{H}_2\text{O}$ , crystals prepared in the manner of Examples 5 and 6 may be dissolved in water to produce a high strength solution.

10 Titanyl sulfate dihydrate filter cake (19 kg) produced using the process described in Example 6 was re-pulped into a pumpable slurry using a solution containing 400 g/L  $\text{H}_2\text{SO}_4$  (4 L) mixed with re-pulp filtrate (36 L) containing 485 g/L free  $\text{H}_2\text{SO}_4$ , 6.7 g/L  $\text{Fe}^{2+}$ , 9.6 g/L  $\text{Fe}^{3+}$  and 5.9 g/L Ti. The slurry was allowed to stir for 15  
15 minutes and then was filtered using a plate and frame filter. A sample of the filtrate from this filtering step was analysed and was found to contain 510 g/L free  $\text{H}_2\text{SO}_4$ , 8.9 g/L  $\text{Fe}^{2+}$ , 10.7 g/L  $\text{Fe}^{3+}$  and 7.4 g/L Ti. Water (50 L) was pumped through the filter to wash the solids. A  
20 sample of the filtrate from the washing step was analysed and found to contain 137 g/L free  $\text{H}_2\text{SO}_4$ , 2.2 g/L  $\text{Fe}^{2+}$ , 3 g/L  $\text{Fe}^{3+}$  and 3.3 g/L Ti. The washed solids were collected and were allowed to dissolve overnight. The resulting titanyl sulfate solution was filtered to remove fine, undissolved  
25 solids, which were predominately silica. The solution was found by assay to contain 467 g/L total  $\text{H}_2\text{SO}_4$ , 1.7 g/L  $\text{Fe}^{2+}$ , 6.5 g/L  $\text{Fe}^{3+}$  and 194 g/L Ti.

## Example 8

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This example describes the conversion of a titanyl sulfate dihydrate slurry directly into a high concentration titanium solution suitable for production of pigment, without an intermediate re-pulp step.

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Titanyl sulfate slurry (108 L) produced from the reactor described in Example 6 was filtered using a

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membrane pressure filter, instead of the belt filter described in Example 6. Recycled filter acid (45 L) containing 338.4 g/L free  $\text{H}_2\text{SO}_4$ , 10.1 g/L  $\text{Fe}^{2+}$ , 2.3 g/L  $\text{Fe}^{3+}$  and 10.1 g/L Ti was mixed with recycled wash water (50 L) containing 93.2 g/L free  $\text{H}_2\text{SO}_4$ , 3.4 g/L  $\text{Fe}^{2+}$ , 0.7 g/L  $\text{Fe}^{3+}$  and 3.4 g/L Ti and with 450 g/L sulfuric acid (10 L). This mixed acid stream was then passed through the membrane pressure filter to wash the filtered solids. The solids were then further washed with water (50 L) and squeezed at a pressure of 4 bar for 5 minutes. Compressed air was then blown through the washed cake for 5 minutes. The filter cake was then removed from the filter and transferred to a container where it dissolved over a period of several hours to give a titanyl sulfate solution containing 218 g/L Ti and 333.5 g/L free  $\text{H}_2\text{SO}_4$ .

#### Example 9

This example describes the precipitation of pigment capable titanium hydroxide from high strength titanyl sulfate solution, using conventional practice.

High strength titanyl sulfate solution (2.5 L) prepared as described in Example 7 was filtered to remove residual solids, then zinc dust (13 g) was added with stirring to remove ferric ions and to generate trivalent titanium. The solution on analysis was found to contain approximately 3.0 g/L of  $\text{Ti}^{3+}$ . Concentrated sulfuric acid was added to give an A/T ratio of  $1.70 \pm 0.05$ . The liquor was then concentrated by evaporation under reduced pressure to give a viscosity of 22-25 cp at  $60^\circ\text{C}$  and  $330 \pm 10$  g/L of  $\text{TiO}_2$  in the final concentrated liquor.

Hydrolysis was carried out based on the Blumenfeld method. A water heel (0.5 L) was heated to  $98 \pm 1^\circ\text{C}$  in a glass reactor equipped with external electrical heating, a temperature controller, thermocouple

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and a rake type stirrer. The pretreated A/T controlled liquor (2.0 L) was separately heated to  $98 \pm 1^\circ \text{C}$  before being added to the water heel at a controlled rate such that all the liquor was added to the heel within  $17 \pm 1$  minutes. The temperature profile was then controlled to precipitate  $\text{TiO}_2$  at a relative rate of 0.7 to 1.0% per minute by ramping the heating rate to give a temperature rise  $0.5^\circ \text{C}$  per min up to the boiling point. Agitation and heating were then stopped for 30 minutes. After this 'stop time' agitation and heating were reapplied to continue precipitation at the rate of 0.7 to 1.0% per minute relative to the initial  $\text{TiO}_2$  concentration. After an overall reaction time of 5 hours the batch was quenched with 2 L of water. Once the solution was cooled to less than  $60^\circ \text{C}$  the solution was vacuum filtered using a Buchner funnel and the precipitate washed with water (6 L) at  $60^\circ \text{C}$ . The cake was allowed to dry by filtration to achieve 30% solids as  $\text{TiO}_2$ . In total 608 g of titanium hydroxide was produced, corresponding to a yield of 96%.

#### Example 10

This example describes the production of rutile seed slurry, which may be used to assist with the rutilisation process during calcination.

Titanium hydroxide filter cake (750 g, loss on ignition 68%) prepared as described in Example 9 was placed in a reaction vessel equipped with agitation and external heating. To the paste, pellets of sodium hydroxide (495 g) were slowly added over 30 minutes. A lid was then placed over the vessel. The temperature was set to  $126^\circ \text{C}$  and was maintained at this level with agitation for a further 60 minutes. At the end of this time the reaction was quenched to  $60^\circ \text{C}$  by adding sufficient water to lower the solids loading to 140 g/L equivalent  $\text{TiO}_2$  (resulting in a total slurry volume of 1713

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mL). The slurry was then filtered using a Buchner funnel, and the precipitate washed with water at 60° C until the wash filtrate contained approximately 1 g/L equivalent Na<sub>2</sub>O, measured using a calibrated conductivity meter.

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The washed filter cake was then transferred to a reflux vessel equipped with an agitator and reslurried to 255 g/L equivalent TiO<sub>2</sub> (giving a slurry volume of 941 mL). The slurry pH was adjusted to 2.8 using concentrated HCl (90 mL, 33% w/v). A 1 g sample was removed to test for cake quality. To the remaining slurry sufficient concentrated HCl (298 mL, 33% w/v) was added to give an HCl:TiO<sub>2</sub> ratio of 0.41, and the temperature was raised to 60° C. The temperature was then increased to the boiling point at a controlled rate of 1° C per minute, and maintained at the boiling point for 90 minutes, after which the slurry was quenched with water to a volume of 2400 mL, giving a solids loading equivalent to 97 g/L TiO<sub>2</sub>. A small sample was neutralized with NaOH, filtered, washed and dried was found by XRD to contain 100% rutile form TiO(OH)<sub>2</sub>.

#### Example 11

This example describes conventional reductive acid leaching of precipitated titanium hydroxide to remove chromophores.

The filtered cake (63.5 g) from Example 9 was slurried in water (0.07 L) in a glass vessel equipped with a laboratory agitator. Concentrated H<sub>2</sub>SO<sub>4</sub> (98%, 9.0 g) was added to the stirred slurry after which coarse rutile nuclei (8.6 mL; prepared as described in Example 10) was added to the slurry to achieve 4% added rutile TiO<sub>2</sub>. The seeded slurry was made up to 0.1 L with water and heated to 75° C. Once at temperature zinc dust was added (0.5 g) and the slurry was maintained at temperature for 2 hours.

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The slurry was then cooled to 60° C and vacuum filtered in a Buchner funnel. The final filtrate was analysed for  $\text{Ti}^{3+}$  concentration to confirm sufficient  $\text{Ti}^{3+}$  was present ( $>0.4$  g/L  $\text{Ti}^{3+}$  preferred (as  $\text{TiO}_2$ )). The cake was then washed  
5 with water at 60° C (three times the volume of precipitate cake). The final cake (60 g) was allowed to dry under vacuum filtration to approximately 30% solids.

#### Example 12

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This example describes calcination of titanium hydroxide to produce a substantially rutilised  $\text{TiO}_2$  calcine with crystal size suitable for pigment production.

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The cake paste (300 g) prepared as described in Example 11 was mechanically mixed in the presence of  $\text{H}_3\text{PO}_4$  (98% solution),  $\text{Al}_2(\text{SO}_4)_3$ ,  $\text{K}_2\text{SO}_4$  to give 0.15%  $\text{P}_2\text{O}_5$ , 0.18%  $\text{Al}_2\text{O}_3$  and 0.28%  $\text{K}_2\text{O}$  as calculated after calcination, until a homogenous mixture is obtained. The paste was the  
20 extruded through a 5 mm die onto glass surface, covered then dried in a 75° C laboratory oven for 12 hours. The solids were then transferred to an electrically heated muffle furnace and the temperature was ramped to 920° C for 3 hours. The calcined solids were removed from the  
25 furnace and allowed to cool to ambient temperature, and the rutilisation measured by XRD was found to be 97.3%.

#### Example 13

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Cooled  $\text{TiO}_2$  solids (800 g) prepared as described in Example 12 were then processed through a laboratory hammer mill and sieved to achieve a particle size of less than 90 microns. The milled particles were then slurried in room temperature water to give a solids loading of 400  
35 g/L (as  $\text{TiO}_2$ ) with the aid of organic dispersant (1,1,1-tris-hydroxymethyl propane). The dispersed slurry was pH adjusted to 10-11 by the addition of 10% w/v NaOH

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solution. The slurry was then passed through a hydraulic bead mill (bead size 0.8-1.0 mm, zirconia stabilized) in recirculation mode until a mean particle size of 0.27  $\mu\text{m}$  was achieved. The slurry was then passed through a 325  $\mu\text{m}$  sieve and the oversize was discarded.

The sieved slurry (2 L) was then transferred to a 3 L beaker and heated to 50° C using an external electric heating mantle. Four solutions (20% w/v  $\text{H}_2\text{SO}_4$ , 10% w/v NaOH, 100 g/L (as  $\text{ZrO}_2$ )  $\text{ZrCl}_2 \cdot 8\text{H}_2\text{O}$  and  $\text{NaAlO}_2$  (caustic stabilized solution containing 17-18% w/w  $\text{Al}_2\text{O}_3$ )) were filled into separate 50 ml burettes and their volumes noted. The reagents were added at temperature such that a final concentration of  $\text{Al}_2\text{O}_3$  (3.5% of  $\text{TiO}_2$  content) and  $\text{ZrO}_2$  (0.88% of  $\text{TiO}_2$  content) was achieved. The slurry was then filtered and washed with water at 60° C to achieve soluble salts in the cake as less than 0.1% as  $\text{Na}_2\text{SO}_4$ , and dried for about 3 hours under vacuum. The cake paste was then mechanically mixed in the presence of organic dispersant to achieve 0.2% carbon (w/w) on the  $\text{TiO}_2$ . The paste was then extruded through a 5 mm die onto glass surface, which was covered and dried in a 75° C laboratory oven for 6 hours to achieve less than 1.0%  $\text{H}_2\text{O}$ . The solids were then lightly hammer milled and the resulting solids passed through a laboratory air microniser which was operated at 6 bar (dried compressed air) for injection and grinding. The micronised product mean particle size was milled to between 0.30 and 0.33  $\mu\text{m}$  as determined by optical density measurements.

#### Example 14

This example shows the ability to continuously hydrolyse high strength titanium solution to produce coarse  $\text{TiO}(\text{OH})_2$  which may be settled and filtered readily.

A continuous pilot plant comprising of 2 x 5 L



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fibre-reinforced plastic (FRP) vessels, equipped with axial turbines and heaters, and an FRP thickener of diameter 30 cm and height 90 cm, equipped with rakes and a rake drive motor, was assembled. The FRP vessels and thickener were arranged in series with cascading overflow pipes between them to allow slurry to flow from vessel to vessel by gravity. An acidic slurry of titanium hydroxide (4 kg) prepared as described in Example 9 was placed in the first vessel as seed, and a solution of 300 g/L of  $\text{H}_2\text{SO}_4$  in water (5 L) was placed in the second vessel to assist the initial start up phase. The vessels were heated to a temperature of 100° C with stirring. On reaching temperature a solution of titanium sulfate prepared as described in Example 7, and containing Ti 130 g/L,  $\text{Ti}^{3+}$  5 g/L, total acid 330 g/L and Fe 10 g/L, was pumped to the first vessel at a rate of 7.5 mL/min. Water was also added at a rate of 6 mL/min to correct for evaporation. On filling of the thickener, a portion of the underflow corresponding to 5 mL/min and 20% w/w solids loading was thereafter continuously pumped to the first vessel to act as seed. In total the hydrolysis pilot plant was operated continuously for 75 hours. On reaching steady state under these process conditions it was found that the vessels and process streams equilibrated to the following compositions.

|               | Ti g/L | $\text{Ti}^{3+}$ g/L | Fe g/L |
|---------------|--------|----------------------|--------|
| Feed solution | 130    | 5                    | 10     |
| Vessel 1      | 70     | 1.4                  | 11     |
| Vessel 2      | 14     | 0.9                  | 9      |

Combined thickener underflow flowrate was 7 mL min (of which 5 mL/min was recycled as described). Equilibrated thickener overflow flowrate was 9 mL/min. The solids loading in the thickener underflow reached 30% w/w by the end of the run. The particle size of the thickener underflow solids was determined using a Malvern

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2000 laser sizer and was found to be  $d_{50}$  7.8  $\mu\text{m}$ .

## CLAIMS

1. A sulfate process for producing titania from a  
titaniferous material (such as ilmenite) of the type which  
5 includes the steps of:

(a) leaching the solid titaniferous material  
with a leach solution containing sulfuric acid and forming  
a process solution that includes an acidic solution of  
10 titanyl sulfate ( $\text{TiOSO}_4$ ) and iron sulfate ( $\text{FeSO}_4$ );

(b) separating the process solution and a  
residual solid phase from the leach step (a);

15 (c) precipitating titanyl sulfate from the  
process solution from step (b);

(d) separating the precipitated titanyl sulfate  
from the process solution;

20 (e) treating the precipitated titanyl sulfate  
and producing a solution containing titanyl sulfate;

(f) hydrolysing the titanyl sulfate in the  
25 solution and forming a solid phase containing hydrated  
titanium oxides and a liquid phase;

(g) separating the solid phase containing  
hydrated titanium oxides and the liquid phase;

30 (h) calcining the solid phase from step (g) and  
forming titania; and

(i) removing iron sulfate from the process  
35 solution from step (b) and/or the depleted process  
solution from step (d).

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2. The process defined in claim 1 further includes supplying the separated process solution from step (d) and/or the separated liquid phase from step (g) to leach step (a).

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3. The process defined in claim 1 or claim 2 includes carrying out titanyl sulfate precipitation step (c) on a continuous basis in a single reactor or in a plurality of reactors in series and/or in parallel.

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4. The process defined in claim 1 or claim 2 includes carrying out titanyl sulfate precipitation step (c) on a batch basis in a single reactor or in a plurality of reactors.

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5. The process defined in claim 3 or claim 4 wherein the reactor or reactors include a vertically extending open-ended draft tube that divides the or reach reactor into an inner chamber (defined by the draft tube) and an outer chamber and a stirring means in the tank.

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6. The process defined in claim 5 wherein titanyl sulfate precipitation step (c) includes supplying process solution containing dissolved titanyl sulfate from leach step (a) and/or removal step (i) to the draft tube reactor or reactors and circulating the process solution successively through the draft tube and the outer chamber of the reactor or reactors for a sufficient period of time to allow precipitation of titanyl sulfate from solution in the process solution.

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7. The process defined in any one of the preceding claims includes controlling titanyl sulfate precipitation step (c) so that the concentration of titanium in the process solution is a relatively low concentration.

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8. The process defined in claim 7 the titanium

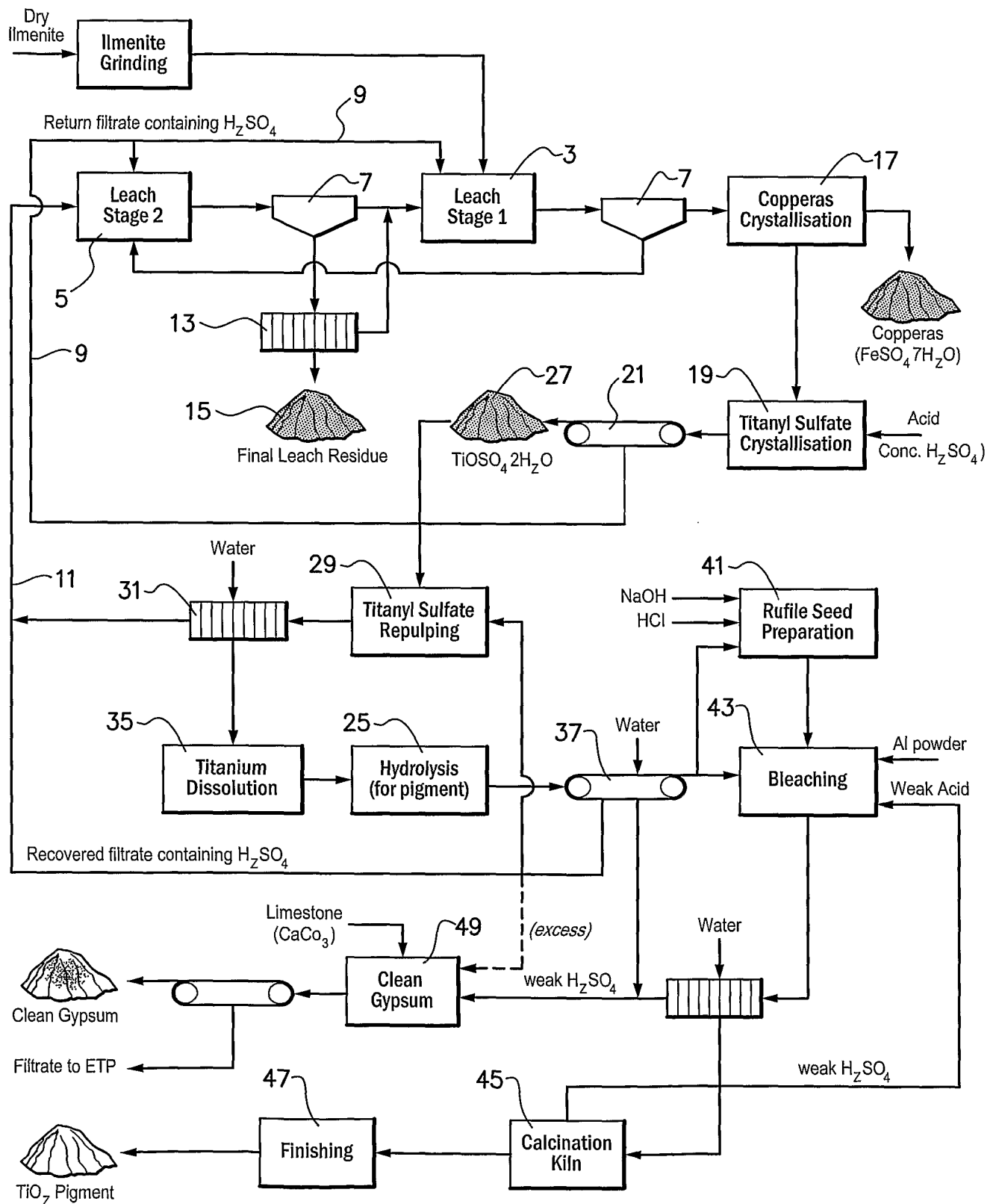
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concentration includes controlling titanyl sulfate precipitation step (c) so that the concentration of titanium in the process solution is less than 25g/L.

5 9. The process defined in any one of the preceding claims includes controlling the process so that there is less than 10 g/L difference between concentration of titanium in the process solution supplied to titanyl sulfate precipitation step (c) and process  
10 solution containing suspended precipitated titanyl sulfate discharged from step (c).

10. The process defined in any one of the preceding claims includes controlling the solids loading in process  
15 solution containing suspended precipitated titanyl sulfate discharged from step (c) be less than 10% solids by weight.

11. The process defined in claim 10 includes  
20 controlling the solids loading in process solution containing suspended precipitated titanyl sulfate discharged from step (c) to be less than 4% solids by weight



## INTERNATIONAL SEARCH REPORT

 International application No.  
**PCT/AU2006/000469**

## A. CLASSIFICATION OF SUBJECT MATTER

Int. Cl.

**C22B 3/08** (2006.01)      **C01G 23/047** (2006.01)      **C22B 3/44** (2006.01)  
**C01G 23/00** (2006.01)      **C22B 3/00** (2006.01)      **C22B 34/12** (2006.01)

According to International Patent Classification (IPC) or to both national classification and IPC

## B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)  
 WPAT; JAPIO; USPTO; Espace@net; IPC (above) and Keywords - titani+, TiO2+, titanyl+, ilmenite+, rutil+, leach+, precipitat+, sul??ric acid, sul??ate+, process+

## C. DOCUMENTS CONSIDERED TO BE RELEVANT

| Category* | Citation of document, with indication, where appropriate, of the relevant passages  | Relevant to claim No. |
|-----------|---|-----------------------|
| P,X       | WO 2005/038060 A (BHP BILLITON INNOVATION PTY LTD), 28 April 2005<br>Whole document | 1-4 & 7-10            |
| X         | WO 2004/035841 A (BHP BILLITON INNOVATION PTY LTD), 29 April 2004<br>Whole document | 1-4, 7 & 10           |
| X         | WO 2004/035842 A (BHP BILLITON INNOVATION PTY LTD), 29 April 2004<br>Whole document | 1-4, 7 & 10           |

☒ Further documents are listed in the continuation of Box C
     
 ☒ See patent family annex

|   |  |
|---|--|
| * Special categories of cited documents:  |  |
| "A" document defining the general state of the art which is not considered to be of particular relevance  | "T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention  |
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| "L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified) | "Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art |
| "O" document referring to an oral disclosure, use, exhibition or other means  | "&" document member of the same patent family  |
| "P" document published prior to the international filing date but later than the priority date claimed  |  |

 Date of the actual completion of the international search  
 16 May 2006

 Date of mailing of the international search report  
 19 MAY 2006

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## INTERNATIONAL SEARCH REPORT

International application No.

PCT/AU2006/000469

| C (Continuation). DOCUMENTS CONSIDERED TO BE RELEVANT |   |                       |
|---|---|-----------------------|
| Category*   | Citation of document, with indication, where appropriate, of the relevant passages          | Relevant to claim No. |
| X   | WO 2004/035843 A (BHP BILLITON INNOVATION PTY LTD), 29 April 2004<br>Whole document         | 1-4, 7 & 10           |
| Y   | CA 2431257 A (ROSSMALI TECHNOLOGY HOLDINGS PTY LTD [ZA]),<br>20 June 2002<br>Whole document | 1-4, 7 & 10           |
| A   | US 4288416 A (DAVIS et al), 8 September 1981<br>Whole document                              | 1-11                  |
| A   | US 4288415 A (RAHM et al), 8 September 1981<br>Whole document                               | 1-11                  |



**INTERNATIONAL SEARCH REPORT**

Information on patent family members

International application No.

**PCT/AU2006/000469**

This Annex lists the known "A" publication level patent family members relating to the patent documents cited in the above-mentioned international search report. The Australian Patent Office is in no way liable for these particulars which are merely given for the purpose of information.

| Patent Document Cited in<br>Search Report |            | Patent Family Member |            |    |            |
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|   |            | ZA                   | 200104774  | WO | 0248412    |
|   |            |                      |            | US | 2004136899 |

Due to data integration issues this family listing may not include 10 digit Australian applications filed since May 2001.

END OF ANNEX